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Structure and properties of Mn₄Cl₉: An antiferromagnetic binary hyperhalogen

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(Received 26 October 2012; accepted 2 January 2013; published online 7 February 2013)

Calculations based on density functional theory show that the structure of Mn₄Cl₉ anion is that of a Mn atom at the core surrounded by three MnCl₃ moieties. Since Mn is predominantly divalent and MnCl₃ is known to be a superhalogen with a vertical detachment energy (VDE) of 5.27 eV, Mn₄Cl₉ can be viewed as a hyperhalogen with the formula unit Mn(MnCl₃)₃. Indeed, the calculated VDE of Mn₄Cl₉ anion, namely 6.76 eV, is larger than that of MnCl₃ anion. More importantly, unlike previously discovered hyperhalogens, Mn₄Cl₉ is the first such hyperhalogen species composed of only two constituent atoms. We further show that Mn₄Cl₉ can be used as a ligand to design molecules with even higher VDEs. For example, Li[Mn(MnCl₃)₂]₂ anion has a VDE of 7.26 eV. These negatively charged clusters are antiferromagnetic with most of the magnetic moments localized at the Mn sites. Our studies show new pathways for creating binary hyperhalogens. © 2013 American Institute of Physics. [http://dx.doi.org/10.1063/1.4776217]

INTRODUCTION

Atoms, molecules, and clusters with high electron affinities (EA) are promising candidates for oxidizing systems with relatively high ionization potentials, and for the synthesis of organic superconductors as well as high energy density salts with unusual stability and desired properties.¹–³ For example, the first compound containing noble gas atom, Xe⁺[PtF₆]⁻, was synthesized by Bartlett using PtF₆, the EA of which was estimated to be about 7.00 eV,⁴ much larger than that of the Cl atom (3.617 ± 0.003 eV).⁶ In 1981, Gusev and Boldyrev coined the term superhalogen to describe such molecules and showed⁷ that these molecules could be created with the formula MXₖ₊₁ where M is a metal atom with maximal valence k and X is a halogen atom.⁷ The vertical detachment energies (VDEs) of MXₖ₊₁ were found to greatly exceed those of the halogen anions due to the delocalization of the additional electron over all the X atoms instead of a single X atom.⁸ Since then, numerous theoretical⁹–¹² as well as experimental¹³–¹⁵ efforts have been devoted to find superhalogen anions with high VDE values. Much of the early work, however, concentrated on having a single metal atom at the core surrounded by halogen atoms. It was later shown that superhalogens can also be formed where the core consisting of a single metal atom is replaced by a metal-halogen complex. Since MXₙ has a closed shell, it is natural to expect that the clusters with the formula of (MXₙ)ₓ will exhibit singlet ground-states. Based on calculations and photoelectron spectroscopy (PES),¹⁶ Boldyrev and Wang found MnₓXₙₓ₊₁ clusters form superhalogens, where M can be an alkaline and alkaline earth atom.³²,³³ In such polynuclear superhalogens, the EA increases with the size of the core. These findings paved the way to design new systems, namely, H₁₂F₁₃ with a very large VDE value of 13.87 eV.³⁴

Most of the earlier studies concentrated on the main group elements of the polynuclear superhalogens. Recently, superhalogens with multiple transition metal atoms at the core are drawing more and more attention. For example, Ta₃F₁₆ was found to have a VDE value of 12.63 eV.²⁰ This cluster can be viewed as (TaF₃)₃F⁻ where the maximal valence of Ta is +5. Similar properties have recently been observed in Mn₃Cl₃ clusters. We note that among all the transition metal elements, Mn plays a unique role in superhalogen studies due to its varying oxidation numbers from 0 to +2 or even negative.³⁵ In particular, Mn exhibits a 3d⁵⁴s² electronic configuration and in the +2 valence state forms the basis of molecular magnets with magnetic moments of 5 μₜ per atom. A search for magnetic superhalogens led to a study of the mass spectra of Mn₃Cl₃ clusters which revealed high peaks for (MnCl₂)Cl⁻ (x = 1–4) composition.³⁶ Theoretical calculations performed for x = 1–3 composition³⁶ showed that these peaks arise due to the anomalously high electron affinities and hence for the enhanced stability of these anions.

Recently, another new class of species was found to have electron affinities even larger than those of superhalogens. These moieties called “hyperhalogens”³⁷–⁴² consist of a central metal atom surrounded by superhalogen moieties. We wondered if for some specific composition polynuclear species may be the same as hyperhalogens. To illustrate this concept we concentrate on Mn₃Cl₃ systems. Since MnCl₃ cluster is a superhalogen one can expect that Mn(MnCl₃)ₓ (x ≥ 3) clusters will behave as hyperhalogens. To find out for what value of x in Mn(MnCl₃)ₓ clusters can also be viewed as a polynuclear species with composition (MnCl₂)ₓCl, we equate:

Mn(MnCl₃)ₓ = (MnCl₂)ₓCl.
This leads to a unique value of $x$, namely, $x = 3$, i.e., Mn$_4$Cl$_9$ can be viewed either as [Mn(MnCl$_3$)$_3$] or (MnCl$_2$)$_4$Cl. In previous work neither theoretical calculations nor experimental measurements of photo-electron spectra were performed for Mn$_4$Cl$_9$. Since the additional electron in Mn$_4$Cl$_9^-$ could not be photodetached with the 6.4 eV laser, it suggests that the electron affinity of Mn$_4$Cl$_9$ must be higher than 6.4 eV. To see if this is indeed the case we studied the geometries and electronic structure of neutral and negatively charged Mn$_4$Cl$_9$. We not only found its electron affinity to be higher than that of Cl, but that it is also higher than that of MnCl$_3$. In addition, the geometry of Mn(MnCl$_3$)$_3$ resembles more of a Mn atom at the center surrounded by three MnCl$_3$ moieties instead of a [MnCl$_5$]$_4$ moiety to which a Cl atom is attached. Thus, Mn$_4$Cl$_9$ can be classified as a hyperhalogen. It is important to note that previously discussed hyperhalogens have all consisted of three elements. Here Mn$_4$Cl$_9$ represents the first binary species having hyperhalogen behavior. In addition, Mn being a transition metal atom that carries a large magnetic moment, these hyperhalogens provide the possibility of designing molecular magnets with large electron affinities.

**COMPUTATIONAL METHODS**

Our calculations were carried out using GAUSSIAN 09 code. Geometry optimization and frequency analysis were performed at B3LYP/6-311+G* level. To search for the ground state of Mn$_4$Cl$_9$ and its anion species, we selected about 20 initial geometrical configurations and optimized each structure without any symmetry constraint. The low-lying isomers ($\Delta E < 5$ kcal/mol) of Mn$_4$Cl$_9$ were then further re-optimized for each possible spin state using the PBE0 hybrid exchange-correlation functional with aug-cc-pVTZ basis set for Cl and SDD basis set and pseudopotential for Mn. For comparison with earlier results, calculations were also performed for Mn$_3$Cl$_2$, Mn$_3$Cl$_4$, Mn$_3$Cl$_6$, Mn$_3$Cl$_8$, and Mn$_3$Cl$_{10}$ for Mn. With earlier results, calculations were also performed for Mn$_4$Cl$_{2k+1}$ ($k = 1 \sim 4$). To validate the consistency of our results with PBE0, VDEs were calculated at the B3LYP/6-311+G*, B3LYP/6-311+G(3df), and OVGF/6-311+G*/PBE0/Mn/SDD/Ci/aug-cc-pVTZ level. For Mn$_4$Cl$_{2k+1}$ ($k = 1, 2$), UHF-CCSD(T)/Mn/SDD/Ci/aug-cc-pVTZ level was used. Zero point energy (ZPE) corrections were calculated by using B3LYP/6-311+G*. We compared the experimental and theoretical VDEs for Mn$_4$Cl$_{2k+1}$ ($k = 1, 2$) to confirm the reliability of our theoretical methods. The natural bond orbital (NBO) analysis was performed for bonding and charge discussions, and GaussView 5.0.8 and VMD 1.9 were used for structure and molecular orbital (MO) visualization.

**RESULTS AND DISCUSSION**

We begin our discussions with the optimized structures of both the neutral and anionic Mn$_4$Cl$_9$ as shown in Figures 1 and 2. Extensive search revealed the trophy-shaped structure (Fig. 1(a)) to be the lowest energy isomer of anionic Mn$_4$Cl$_9^-$. On the other hand, the ground-state geometry of neutral Mn$_4$Cl$_9$ completely differs from that of its anion displaying an open, linear structure (Fig. 2(a)). The ground-state geometries for both neutral and anionic Mn$_4$Cl$_9$ have three Mn atoms with fourfold coordination while the fourth Mn atom is threefold coordinated.

To refine the geometries of the isomers, optimizations at PBE0/Mn/SDD/Cl/aug-cc-pVTZ were carried out. Three energetically degenerate structures were found as shown in Figure 3, where the trophy-shaped structure (Fig. 3(a)) is again the lowest-energy configuration; the second isomer has nearly the same energy ($\Delta E = 0.01$ eV) and the third one is only 0.06 eV higher in energy as compared to the first one. The magnetic couplings between Mn atoms in all the three structures are antiferromagnetic, resulting in the total magnetic moments of 0 $\mu_B$. The ground-state structure of anionic Mn$_4$Cl$_9$ has C$_1$ symmetry, where Mn-Mn distance is different for the different spin orientations. The low-lying isomer (Figure 3(c)) has C$_2$ symmetry having a diamond-like shape. Based on our calculations, we found that the smaller 6-311G basis sets would underestimate the spin magnetic moments. Therefore, we used PBE0/Mn/SDD/Cl/aug-cc-pVTZ for magnetic moment and orbital analyses.

The Mulliken atomic spin density distribution for the ground-state structures of anionic Mn$_4$Cl$_9$ as well as neutral Mn$_4$Cl$_9$ are given in Figure 4, indicating that the majority of the spin moments reside on the four Mn atoms with fourfold coordination.
atoms. As stated above, the spin coupling in anionic Mn₄Cl⁹ is antiferromagnetic, where the four Mn atoms carry magnetic moments of 5.00 μₜ, 5.12 μₜ, 5.03 μₜ, and 5.02 μₜ, respectively. We note that in Mn₄Cl⁹, the corresponding α spin electron configurations of the four Mn atoms are [Ar]4s⁰.19 3d⁰.94 4p⁰.21, [Ar]4s⁰.14 3d⁰.43 4p⁰.28, [Ar]4s⁰.18 3d⁰.91 4p⁰.27, and [Ar]4s⁰.11 3d⁰.35 4p⁰.23, while the β spin electron configurations of the four Mn atoms are [Ar]4s⁰.13 3d⁰.29 4p⁰.19, [Ar]4s⁰.17 3d⁰.85 4p⁰.30, [Ar]4s⁰.15 3d⁰.35 4p⁰.25, and [Ar]4s⁰.19 3d⁰.91 4p⁰.28, respectively. It is therefore clear that although the main contributions to the moments come from 3d orbitals, due to the s-p-d hybridization, the 4s and 4p orbitals are slightly polarized, resulting in large total magnetic moments on Mn sites. However, when one electron is removed from the anion, the s-p-d hybridization is reduced. Consequently, the spin couplings between Mn sites become ferrimagnetic with the total magnetic moment of 1 μₜ and the corresponding moments at the four Mn sites are 5.00 μₜ, −5.06 μₜ, 5.03 μₜ, and −4.30 μₜ respectively. The relative energies of Mn₄Cl⁹ for different possible spin states calculated with respect to the lowest-energy spin configuration are shown in Figure 5. The doublet spin state with three AFM coupled Mn pairs is found to be the energy minimum, which is analogous to the results obtained from the previous analysis of Mn₂Cl⁵ and Mn₃Cl⁷.³⁶

From the above calculations we can see that the anionic Mn₄Cl⁹ is an antiferromagnetic molecular magnet. To have further insight into its stability and the oxidation property of its neutral counterpart, we used B3LYP and PBE0 hybrid functional with various basis sets to calculate the VDEs of anionic Mn₄Cl⁹ (x = 1 - 4) clusters. Outer valence Green’s function (OVGF) and coupled cluster with single and double and perturbative triple excitations (CCSD(T)) methods were also employed for comparisons. Theoretical VDEs obtained at different levels of theory are compared with experimental values of Mn₄Cl₂⁺ (x = 1 - 2) in Table I. Just as expected, the VDEs of anionic Mn₄Cl₂⁺ (x = 1 - 4) increase with the value x. For x = 1, the calculated VDE is 5.27 eV at the PBE0/MnSDD/Cl/aug-cc-pvtz level, whereas the calculated VDE for x = 4 is found to increase to 6.76 eV. Unfortunately, this cannot be compared with the previous experiment due to the limitation of the PES photon energy used.²⁹,³⁶ Our theoretical results for x = 1 and x = 2, however, are in good agreement with previous calculations as well as experimental PES results,³⁶ suggesting that it is reasonable to employ DFT method within hybrid density functional to acquire theoretical VDE for anionic Mn₄Cl⁹. As can be seen, the VDEs for all the anionic Mn₄Cl₂⁺ (x = 1 - 4) species greatly exceed the EA of chlorine atom and thus may be classified as superhalogens. However, they do not belong to conventional superhalogens and can be considered as (MnCl₂)₂Cl as is pointed out in previous effort.³⁶
TABLE I. Comparison between experimental and calculated VDEs values of $\text{Mn}_k\text{Cl}_{2k+1}$ ($k = 1 \sim 4$). All energies are in eV.

<table>
<thead>
<tr>
<th></th>
<th>B3LYP/6-311G$^*$</th>
<th>B3LYP/6-311+G(3df)</th>
<th>PBE0/Cl/aug-cc-pVTZ</th>
<th>PBE0/Mn/SDD/Cl/aug-cc-pVTZ</th>
<th>OVGF$^a$</th>
<th>UHF-CCSD(T)$^b$</th>
<th>ROHF-CCSD(T)$^c$</th>
<th>EOMCCSD$^d$</th>
</tr>
</thead>
<tbody>
<tr>
<td>$\text{MnCl}_3$</td>
<td>5.18</td>
<td>5.05</td>
<td>5.31</td>
<td>5.27(5.6±0.1)$^f$</td>
<td>5.65</td>
<td>5.83</td>
<td>5.68</td>
<td>5.79</td>
</tr>
<tr>
<td>$\text{Mn}_2\text{Cl}_5$</td>
<td>6.10</td>
<td>5.98</td>
<td>6.24</td>
<td>6.21(6.3±0.1)$^f$</td>
<td>6.69</td>
<td>6.76</td>
<td>f</td>
<td>f</td>
</tr>
<tr>
<td>$\text{Mn}_3\text{Cl}_7$</td>
<td>6.37</td>
<td>6.25</td>
<td>6.54</td>
<td>6.51</td>
<td>7.18</td>
<td>f</td>
<td>f</td>
<td>f</td>
</tr>
<tr>
<td>$\text{Mn}_4\text{Cl}_9$</td>
<td>6.61</td>
<td>6.54</td>
<td>6.77</td>
<td>6.76</td>
<td>7.24</td>
<td>f</td>
<td>f</td>
<td>f</td>
</tr>
</tbody>
</table>

$^a$VDEs were calculated at OVGF/6-311+G(3df)/PBE0/Mn/SDD/Cl/aug-cc-pvtz level.
$^b$VDEs were calculated at UHF-CCSD(T)/Mn/SDD/Cl/aug-cc-pvtz/PBE0/Mn/SDD/Cl/aug-cc-pvtz level.
$^c$VDEs were calculated at ROHF-CCSD(T)/Mn/SDD/Cl/aug-cc-pvtz/PBE0/Mn/SDD/Cl/aug-cc-pvtz level.
$^d$VDEs were calculated at EOMCCSD/Mn/SDD/Cl/aug-cc-pvtz/PBE0/Mn/SDD/Cl/aug-cc-pvtz level.
$^e$VDEs in parentheses are experimental values.$^{36}$
$^f$VDE could not be calculated at this level of theory.

That the VDEs of polynuclear superhalogens become larger with the increasing core size has inspired growing interest in the search of superhalogens with extremely high VDEs for oxidation of materials with large ionization potential. Hyperhalogens also provide a means of creating species with ever increasing electron affinities, but a link between these two concepts has not yet been demonstrated. Here we present a structural-based hypothesis to elucidate the unconventionally high VDE for anionic $\text{Mn}_4\text{Cl}_9$. $\text{MnCl}_3$ is a typical high spin superhalogen since the preferred oxidation state of Mn in Mn halides is $+2$. If we take $\text{MnCl}_3$ as a structural block, $\text{Mn}_2\text{Cl}_5$ can be understood as $\text{Mn}(\text{MnCl}_3)\text{Cl}_2$, $\text{Mn}_3\text{Cl}_7$ as $\text{Mn}(\text{MnCl}_3)_2\text{Cl}$, and $\text{Mn}_4\text{Cl}_9$ as $\text{Mn}(\text{MnCl}_3)_3$. In each of these clusters the core is a Mn atom and halogen (Cl) atoms are successively replaced by $\text{MnCl}_3$ superhalogen moieties. Thus, it is expected that the VDE of $\text{Mn}_k\text{Cl}_{2k+1}$ ($k = 1 \sim 4$) increases as the number of Mn atoms getting larger. All these results except for $\text{Mn}_4\text{Cl}_9$ have been previously confirmed. If the above analysis will hold for $\text{Mn}_4\text{Cl}_9$ it can be classified as a hyperhalogen composed of only two elements. We define such kind of hyperhalogens as binary hyperhalogen because there is only one element acting as both the central atom of the hyperhalogen and the central atom of the superhalogen moieties.

It is therefore appropriate to presume that the ground-state geometry for anionic $\text{Mn}_4\text{Cl}_9$ involves a central Mn atom chelated by three $\text{MnCl}_3$ blocks with every $\text{MnCl}_3$ fragment providing two Cl bridging atoms, forming a highly symmetric octahedral structure. However, based on B3LYP calculation, the octahedral form (Fig. 1(f)) was found to be higher in energy than the trophy-shaped framework (Fig. 1(a)) by 0.66 eV. We attribute this to an intrinsic property of central metal cation, namely, coordination number. Mn$^{2+}$ is most likely to display the coordination number of 4 in its halides,$^{35}$ thus preferring the $sp^3$ hybridization as well as tetragonal geometry. Our calculations show that the global minimum structure for the anionic $\text{Mn}_4\text{Cl}_9$ is determined by its hyperhalogen...
nature as well as the coordination number of Mn$^{2+}$. Thus, the Mn$_4$Cl$_9$ is different from the conventional hyperhalogens.

In order to have a better understanding of the nature of bonding and electronic structure in the neutral and negatively charged Mn$_4$Cl$_9$, we calculated the electron distribution by analyzing the NBO charge. The results for neutral and anionic Mn$_4$Cl$_9$ are presented in Figure 6. We can clearly see that the charges on all the Mn atoms are positive while they are negative on all the Cl atoms, resulting from the anticipated Mn to Cl charge transfer. However, owing to partial back electron transfer from Mn to Cl, the Mn–Cl bonds have a somewhat covalent nature. This is confirmed by the relatively smaller charges on Mn and Cl atoms compared to other metal halides. The Cl atoms are, therefore, slightly polarized and also carry small spin magnetic moments (Figure 4). Particularly, the central Mn atom in the anionic Mn$_4$Cl$_9$ and the fourfold coordinated terminal Mn atom in the neutral Mn$_4$Cl$_9$ involve more of their 3$d$ electrons in the Mn–Cl interactions, as it can be seen from the smaller NBO charges compared to other Mn atoms. The total charges on all the Mn atoms in the negatively charged Mn$_4$Cl$_9$ and its neutral counterpart are +3.44 e and +3.31 e, respectively. This clearly indicates that the extra electron in the anionic Mn$_4$Cl$_9$ is delocalized over the negatively charged Mn$_4$Cl$_9$ and its neutral counterpart are to other Mn atoms. The total charges on all the Mn atoms in as it can be seen from the smaller NBO charges compared to other metal halides. The Cl atoms are, therefore, slightly polarized and also carry small spin magnetic moments (Figure 4). Particularly, the central Mn atom in the anionic Mn$_4$Cl$_9$ and the fourfold coordinated terminal Mn atom in the neutral Mn$_4$Cl$_9$ involve more of their 3$d$ electrons in the Mn–Cl interactions, as it can be seen from the smaller NBO charges compared to other Mn atoms. The total charges on all the Mn atoms in the negatively charged Mn$_4$Cl$_9$ and its neutral counterpart are +3.44 e and +3.31 e, respectively. This clearly indicates that the extra electron in the anionic Mn$_4$Cl$_9$ is delocalized over the Cl atoms, which contributes considerably to the high VDE of anionic Mn$_4$Cl$_9$. MO analysis indicates that the highest occupied $\alpha$ and $\beta$ spin orbitals of anionic Mn$_4$Cl$_9$ are energetically nearly degenerate and mainly consist of the terminal Cl $p_z$ orbital and the adjacent Mn $d_{x^2−y^2}$ orbital, while the lowest unoccupied $\beta$ spin orbital is mostly contributed by the fourfold coordinated terminal Mn $d_{xz}$ orbital and the peripheral Cl $p_x$ orbitals, as illustrated in Figure 7.

To further explore the possibility that the Mn$_4$Cl$_9$ can serve as a building block for larger clusters, we calculated the equilibrium structure and VDE of anionic Li(Mn$_4$Cl$_9$)$_2$ cluster. The ground state geometry of anionic Li(Mn$_4$Cl$_9$)$_2$ cluster (Fig. 8) is linear, where bond lengths of the two Mn$_4$Cl$_9$ fragments are nearly identical with those of the free anionic Mn$_4$Cl$_9$, indicating that the Mn$_4$Cl$_9$ anion is highly stable. We then compared the VDE of anionic Li(Mn$_4$Cl$_9$)$_2$ with several other superhalogen and hyperhalogen anions to find out whether the anionic Li(Mn$_4$Cl$_9$)$_2$ exhibited an ultra-hyperhalogen nature. The results are given in Table II. At the same level of theory, Li(Mn$_4$Cl$_9$)$_2$ anion possesses a high VDE of 7.26 eV which is much larger than that of LiF$_2$ or Li(BO$_2$)$_2$ anion. Furthermore, the VDE of Li(Mn$_4$Cl$_9$)$_2$ anion even exceeds that of the free Mn$_4$Cl$_9$ anion, which supports the conclusion that the Li(Mn$_4$Cl$_9$)$_2$ is indeed an ultra-hyperhalogen. From the calculated VDEs in Table II, it is also obvious that the VDEs of the LiX$_2^−$ clusters increase as the ligands having higher VDEs. It is thus worth employing superhalogens or hyperhalogens as building blocks of clusters with unprecedented oxidizing property. The fragmentation energy of Li(Mn$_4$Cl$_9$)$_2$ corresponding to the lowest energy pathway is 11.56 eV. Consequently, anionic Li(Mn$_4$Cl$_9$)$_2$ is thermodynamically stable. Thus, it should be possible to synthesize these in experiments.

### CONCLUSIONS

In summary, we have systematically studied the equilibrium geometries, magnetic couplings, vertical detachment energies, and the charge distributions of Mn$_4$Cl$_9$ and its anion by using density functional theory based calculations. The following conclusions are drawn: (1) There exists a unique composition of Mn$_4$Cl$_9$ where the cluster can be simultaneously termed as a binary hyperhalogen or a polynuclear superhalogen. In this case Mn$_4$Cl$_9$ is such a cluster. (2) The ground state geometry of Mn$_4$Cl$_9$ anion can be viewed as Mn(MnCl$_3$)$_3$ with some distortions. Since the MnCl$_3$ behaves like a magnetic superhalogen with Mn atom carrying a magnetic moment of 5 $\mu_B$, Mn$_4$Cl$_9$ anion can be called a magnetic hyperhalogen. This is a unique binary system since previously studied hyperhalogens consisted of three or more elements like Au(BO$_2$)$_2$ and were not magnetic. (3) In the anionic Mn$_4$Cl$_9$, the magnetic couplings between Mn atoms are antiferromagnetic and each Mn site carries a moment of about 5 $\mu_B$. Cl sites are slightly spin polarized, sharing similar features with Mn$_2$Cl$_3$ and Mn$_3$Cl$_7$. (4) Natural bonding orbital analysis demonstrates the partial covalent nature of Mn–Cl interactions. (5) The Mn$_4$Cl$_9$ can be used as a ligand to design ultra-hyperhalogen moieties. By varying the ligand X from F to BO$_2$ to Mn$_4$Cl$_9$, the VDE of LiX$_2^−$ can be seen to steadily increase from 5.44 eV to 5.80 eV to 7.26 eV. We hope that our study will stimulate more experimental efforts to better understand binary hyperhalogens with varying size.

### Table II. Comparison of VDEs for X$^−$ and LiX$_2^−$ (X = F, BO$_2$, Mn$_4$Cl$_9$).

<table>
<thead>
<tr>
<th>Anions</th>
<th>F$^−$</th>
<th>BO$_2^−$</th>
<th>Mn$_4$Cl$_9$</th>
<th>LiF$^−$</th>
<th>LiBO$_2^−$</th>
<th>Li(BO$_2$)$_2^−$</th>
<th>Li(Mn$_4$Cl$_9$)$_2^−$</th>
</tr>
</thead>
<tbody>
<tr>
<td>VDE (eV)</td>
<td>3.65</td>
<td>4.14</td>
<td>6.76</td>
<td>5.44</td>
<td>5.66</td>
<td>5.80</td>
<td>7.26</td>
</tr>
</tbody>
</table>

### FIG. 8. Optimized geometries of Li(Mn$_4$Cl$_9$)$_2$(a) and Mn$_4$Cl$_9$(b). All bond lengths are in Å.
ACKNOWLEDGMENTS

This work is supported by grants from the National Natural Science Foundation of China (Grant Nos. NSFC-21273012 and NSFC-11174014), the National Grand Fundamental Research 973 Program of China (Grant No. 2012CB921404), and from the (U.S.) Department of Energy.

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